

Date Planned ://_ Actual Date of Attempt ://_					Daily Tutorial Sheet-2			Expected Duration : 90 Min	
					Level-1		Exact Duration :		
16.	Extra pure N_2 can be obtained by heating:							\odot	
	(A)	NH ₃ with C	CuO		(B)	$\mathrm{NH_4NO_3}$			
	(C)	$(NH_4)_2Cr_2C$) ₇		(D)	$Ba(N_3)_2$			
17.	Which of the following halogen oxide is used for estimation of carbon monoxide in automobile exhibits gases?								
	(A)	${\rm Cl_2O_7}$	(B)	$\rm I_2O_5$	(C)	${ m ClO}_2$	(D)	BrO_3	
18.	An explosive compound (A) reacts with water to produce $\mathrm{NH_4OH}$ and $\mathrm{HOCl}.$ Then, the compound (A), is :								
	(A)	TNG	(B)	NCl_3	(C)	PCl_3	(D)	HNO_3	
19.	The atomicity of phosphorus is X and the PPP bond angle in the molecule is Y. What are X and Y?								
	(A)	$X = 4, Y = 90^{\circ}$			(B)	$X = 4, Y = 60^{\circ}$			
	(C)	$X = 3, Y = 120^{\circ}$			(D)	$X = 2, Y = 180^{\circ}$			
20.	H_3PO_3 has Non-ionisable $P-H$ bonds.								
	(A)	3	(B)	1	(C)	2	(D)	None of these	
21.	Which is tribasic acid?								
	(A)	H_3PO_2	(B)	H_3PO_4	(C)	$\mathrm{H_4P_2O_7}$	(D)	H_3PO_3	
22.	The c	The catalyst used in manufacture of ammonia is :							
	(A)	V_2O_5	(B)	Pt	(C)	Fe	(D)	Ni(CO) ₄	
23.	What are the products obtained when ammonia is reacted with excess chlorine?								
	(A)	(A) N ₂ and NCl ₃			(B)	${\rm N_2}$ and HCl			

24. Zinc and cold dil. HNO_3 reacts to produce :

 $\ensuremath{\text{N}}_2$ and $\ensuremath{\text{NH}}_4\ensuremath{\text{Cl}}$

(A) NO

(B) NO_2

(C) NH_4NO_3

 $\ensuremath{\mathsf{NCl}}_3$ and $\ensuremath{\mathsf{HCl}}$

(D) $Zn(OH)_2$

25. Pnicogens are the elements of group:

(A)

(C)

(B) 13

(C) 8

(D) Zero

26. The treatment Cu with dilute HNO_3 gives :

(A)

 N_2O

15

(B) NO

(C)

(D)

 NH_4^+

(D) NO₂

27. Which of the following oxy-acids of phosphorus is a reducing agent and monobasic?

 $(A) \qquad H_3PO_2$

(B)

 H_3PO_3

(C)

 H_3PO_4

(D)

 $H_4P_2O_6$

28. Pb reacts with dilute HNO_3 gives :

(A) NO

(B) NH_4NO_3

(C)

 N_2O_5

(D) NO_2



29. Calcium cyanamide on treatment with steam produces :

(A)
$$NH_3 + CaO$$

(B)
$$NH_3 + CaHCO_3$$

(C)
$$NH_3 + CaCO_3$$

(D)
$$NH_3 + Ca(OH)_2$$

30. $CH_2 < COOH \xrightarrow{P_4O_{10}. 150^{\circ}C} X$



Compound (X) is:

(A) malonic acid

(B) carbon suboxide

(C) tartaric acid

(D) acetic acid